



# IJEAST

INTERNATIONAL JOURNAL  
OF ENGINEERING APPLIED SCIENCE  
AND TECHNOLOGY



**VOLUME : 5    ISSUE : 1    Print / Issue Publication Date: 13-Jul-2020**



**ISSN : 2455-2143**



**DOI : 10.33564/IJEAST.2020.v05i01.087**

Indexed In



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# REMOVAL OF IRON FROM NATURAL GROUND WATER BY USING LOCALLY AVAILABLE MATERIAL

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**Abstract— Adsorption, being a simple, practical and effortless procedure, is an effective procedure for the expulsion of toxins from wastewater. This study mainly deals with the removal of iron from ground water sources by using locally available low cost materials such as rice husk. The sample used for this filtration process was synthetically prepared iron solution using ferrous ammonium sulphate. In this examination rice husk have appeared to be successful and minimal effort adsorbent. Evacuation of iron from watery arrangements was contemplated utilizing bounteously accessible rice husk under different trial conditions. The most expulsion effectiveness of 96% was accomplished at rice husk of depth 2cm as filter medium for iron (II) removal at equilibrium conditions. Adsorption qualities showed that rice husk were acceptable for the purpose.**

Keywords— Rice husk, Iron content, Absorbent

## I. INTRODUCTION

Areas having a tropical climate have a high concentration of iron (Fe) in groundwater. In groundwater, Fe generally occurs in the oxidation state reduced soluble divalent ferrous iron ( $\text{Fe}^{2+}$ ). When groundwater comes in contact with oxygen of the atmosphere, the Fe is oxidized to the ferric state and is precipitated as Fe-mineral. The subsurface reducing conditions have significant influence on the high Fe content of groundwater. Iron in rural groundwater supplies is a common problem, its concentration level ranges from 0 to 50 mg/l, while WHO recommended level is  $< 0.3$  mg/l. When iron content exceed this permissible value then it can causes hemochromatosis, stomach problems, nausea, vomiting, wrinkles in skin. There are various methods of removal of dissolved iron from groundwater, some of them are Aeration or gas transfer, Adsorption, Ion exchange, Electro chemical, Desalination.

## II. METHODOLOGY

### Preparation of synthetic iron water

For the synthesis of synthetic iron water sample, Ferrous ammonium sulphate has been used in varying amounts in order to produce the required concentration for the tests.

### Preparation of adsorbent

**Rice Husk:** The grounded rice husk is collected from locally available mill. It is prepared by washing it with tap water followed by distilled water. It is then allowed to dry for 3 days in the sun. The dried rice husk is then sieved to get size smaller than 300 micrometer. Then it was ready to use as an adsorbent.

### Setting up of filter

#### Filter containing charcoal as adsorbent:

- The filter is of height 27 cm and diameter 22cm.
- In the first layer, coarse aggregates of sizes passing through IS sieve 40mm and retained on 20mm are used at thickness of 3cm.
- The second layer of coarse aggregates used are of sizes passing through IS sieve 10mm and retained on 4.75mm at thickness 2cm.
- The third layer is of coarse sand of size passing through IS sieve 4.75mm and retained on 2mm at thickness of 1cm.
- The fourth layer is of medium sand of size passing through IS sieve 2mm and retained on 600 micron at thickness 4cm.
- The fifth layer is of charcoal which is of granular type of approximate size between 40mm and 4.75mm of thickness 5cm.
- The sixth layer is again of medium sand of thickness 4 cm.
- A layer of muslin cloth has been used above the charcoal layer to prevent it from mixing with sand.

#### Filter containing agricultural adsorbent:

- The filter is of height 33 cm and diameter 25cm. The layers of filter media are laid using sand, coarse aggregates and the appropriate adsorbent.
- In the first layer, coarse aggregates used are of sizes passing through IS sieve 40mm and retained on 20mm at thickness of 5cm.



- The second layer of coarse aggregates used are of sizes passing through IS sieve 10mm and retained on 4.75mm at thickness 3cm.
- The third layer is of coarse sand of size passing through IS sieve 4.75mm and retained on 2mm at thickness of 4.5cm.
- The fourth layer is of medium sand of size passing through IS sieve 2mm and retained on 600 micron at thickness 8cm.
- The fifth layer is of the agricultural adsorbent, 1cm for rice husk.
- The sixth layer is of medium sand of above mentioned size of thickness 6 cm.
- A layer of muslin cloth is used once below the adsorbent layer and once above it to prevent the merging of the individual layers.

After allowing the water to pass through both the filter medium, ferrous iron, Fe (II) content in the water sample is estimated by titrating using potassium permanganate solution.

**Procedure:**

- Approximately 0.8g of  $KMnO_4$  is weighed and transferred to 250ml water taken in a 500ml beaker and dissolved. The solution is boiled gently for about 20 minutes and then cooled to room temperature. The solution is filtered and stored in a clean dark brown coloured bottle.
- About 0.63g of oxalic acid is weighed out accurately and transferred into a 10ml volumetric flask and dissolved in water. Make up the volume to 100ml mark.
- 10ml of 0.1N oxalic acid is pipetted out into a 250ml conical flask. 10ml of 4N  $H_2SO_4$  is added and mixed well. The solution is heated on wire gauze till the flask is unbearable to touch ( $-60^\circ C$ ). Now it is titrated with  $KMnO_4$  giving swirling motion to the flask. The end point will be a faint pinkish colour. The process is repeated for concordant readings.

**III. RESULTS AND DISCUSSIONS**

**Observations by titration for the analysis of iron in water (not suitable for low concentration solutions)**

Let, Strength of Oxalic acid solution =  $N_1$   
 Volume of Oxalic acid solution,  $V_1=10ml$   
 Strength of potassium permanganate solution ( $KMnO_4$ )=  $N_2$   
 Volume of potassium permanganate solution ( $KMnO_4$ ) =  $V_2$   
 Burette readings:

Table 1: Standardisation of  $KMnO_4$  solution

Sl no.	Initial reading (ml)	Final reading (ml)	Volume of $KMnO_4$ $V_2$ (ml)	Concordant Reading for $V_2$ (ml)
1	0	8.9	8.9	9.6
2	8.9	18.9	10	

3	18.9	29	10.1	
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**Calculation:**

10ml of 0.1N oxalic acid solution =  $X$  ml of  $KMnO_4$  solution  
 We have,

$$V_1 * N_1 = V_2 * N_2$$

$$N_2 = (V_1 * N_1) / V_2$$

$$= (10 * 0.1) / 9.6$$

$$= 0.1041 N$$

Table 2: Burette readings: (For iron sample before passing through charcoal filter)

Sl no.	Initial readings (ml)	Final readings (ml)	Volume of $KMnO_4$ (ml)	Concordant reading for $V_3$ (ml)
1	0	15.8	15.8	12.63
2	15.8	26.9	11.3	
3	26.9	37.9	11	

Table 3: Burette readings (For iron sample after passing through charcoal filter)

Sl no.	Initial readings (ml)	Final readings (ml)	Volume of $KMnO_4$ (ml)	Concordant reading for $V_4$ (ml)
1	0	2.3	2.3	2.13
2	2.3	4.3	2	
3	4.3	6.4	2.1	

**Calculations (for charcoal filter):**

1ml of 0.1N  $KMnO_4$  = 0.005585g, Fe  
 $V_3$  ml of  $N_2$   $KMnO_4$  =  $(0.005585 * V_3 * N_2) / 0.1$   
 = 0.0734g, Fe  
 10ml of  $Fe^{2+}$  solution contains = 0.0734g, Fe  
 Amount of Fe per litre of the given solution =  $0.0734 * 100g$   
 = 7.34g

Similarly  $V_4$  ml of  $N_2$   $KMnO_4$  = 0.01238g, Fe  
 Amount of Fe per litre of the given solution = 1.23g

Table 4: Burette readings (For iron sample before passing through rice husk filter)

Sl. No.	Initial readings (ml)	Final Readings (ml)	Volume of $KMnO_4$ (ml)	Concordant reading for $V_3$ (ml)
1.	0	15	15	12.3
2	15	26.3	11.3	
3	26.3	36.9	10.6	



Table 5: Burette readings (For iron sample after passing through rice husk of 1cm depth filter)

Sl. No.	Initial readings (ml)	Final Readings (ml)	Volume of KMnO <sub>4</sub> (ml)	Concordant reading for V4 (ml)
1.	0	1.2	1.2	1.16
2	1.2	2.4	1.2	
3	2.4	3.5	1.1	

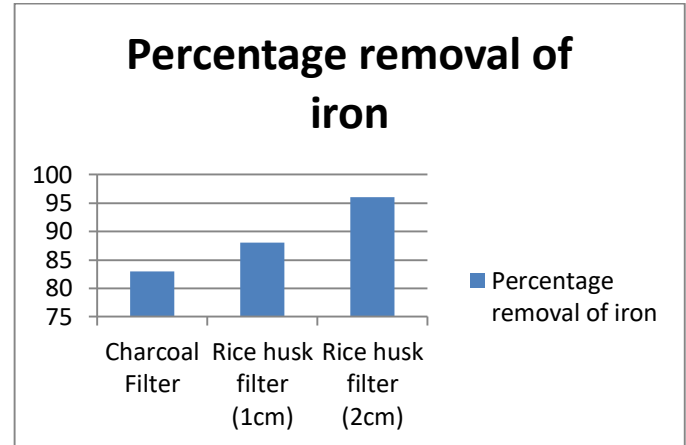


Fig 1: Comparison between the charcoal filter and rice husk filter

Table 6: Burette readings (For iron sample before passing through rice husk filter)

Sl. No.	Initial readings (ml)	Final Readings (ml)	Volume of KMnO <sub>4</sub> (ml)	Concordant reading for V5 (ml)
1.	0	15.1	15.1	12.63
2	15.1	26.7	11.6	
3	26.7	37.9	11.2	

#### IV. CONCLUSION

From the study it was found that removal of iron was found to be 83% for charcoal filter and 88% and 96% in case of rice husk used filter of depth 1cm and 2cm respectively. Therefore with increase in the adsorbent dosage leads to increase in the removal of iron content from the sample.

Again Rice husk has been found to be in reasonable priced and an effective adsorbent for the elimination of Fe (II) ions from aqueous solution without requiring any pre-treatment. The test outcome has proven to remove iron to a maximum limit.

#### V. ACKNOWLEDGEMENT

At the very outset, we are pleased and highly honoured to express our sincere and heartfelt gratitude to Miss Jeba Sahana, Miss Ishani Baishya, Mr Shibaraj Brahma Gayari, Mr Rohit Singha, Students of Department of Civil Engineering, Girijananda Chowdhury Institute of Management & Technology, Hathkhowapara, Azara, Guwahati-17, for their constant support, inspiration and full co-operation throughout the project work.

We would also like to take the opportunity to thank Prof. (Dr.) Krishnanga Gohain, Head of Department, Civil Engineering & Technology, Girijananda Chowdhury Institute of Management & Technology, Hathkhowapara, Azara, Guwahati-17.

Also we are gratified towards all the faculties and staff members of the Department of Civil Engineering, Chowdhury Institute of Management & Technology, for their support to make the project work a reality.

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Table 7: Burette readings (For iron sample after passing through rice husk of 2cm depth filter)

Sl. No.	Initial readings (ml)	Final Readings (ml)	Volume of KMnO <sub>4</sub> (ml)	Concordant reading for V6 (ml)
1.	0	0.5	0.5	0.4
2	0.5	1.1	0.6	
3	1.1	1.3	0.2	

#### Calculations (for filter containing rice husk):

- 1ml of 0.1N KMnO<sub>4</sub> = 0.005585g, Fe
- V<sub>3</sub> ml of N<sub>2</sub> KMnO<sub>4</sub> = (0.005585 \* V<sub>3</sub> \* N<sub>2</sub>) / 0.1  
= 0.0715g, Fe
- 10ml of Fe<sup>2+</sup> solution contains = 0.0715g, Fe
- Amount of Fe per litre of the given solution = 0.0715 \* 100g = 7.15g
- Similarly V<sub>4</sub> ml of N<sub>2</sub> KmnO<sub>4</sub> = 0.00674g, Fe
- Amount of Fe per litre of the given solution = 0.68g
- Similarly V<sub>5</sub> ml of N<sub>2</sub> KmnO<sub>4</sub> = 0.0734g, Fe
- Amount of Fe per litre of the given solution = 7.35g
- Similarly V<sub>6</sub> ml of N<sub>2</sub> KmnO<sub>4</sub> = 0.0023g, Fe
- Amount of Fe per litre of the given solution = 0.23g



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2455-2143